• *Local Section News*

Northern California Section

The Northern California Section of the AOCS held its spring meeting Friday, May 18 at the Claremont Hotel in Berkeley. Following the social hour and dinner, Chairman L. A. Glodblatt introduced the two speakers: C. E. Claassen, President, Pacific Oilseeds, Inc., of Woodland, Calif., and Lowell Cummings, Chemist, Pacific Vegetable Oil Corp., Richmond, Calif.

Dr. Claassen reviewed the history and development of safflower as a commercial crop in the U. S. Mr. Cummings spoke on the companion story of developing the industrial utilization of safflower oil, particularly in surface coatings.

The next Northern California Section meeting is slated for Fresno's Towne and Country Lodge for Friday evening and Saturday morning, September 21–22. The featured speaker for Friday evening is R. B. Alfin-Slater of the University of California School of Public Health, Los Angeles, who will speak on "How to Succeed in Reducing Without Counting Calories." Two technical presentations are scheduled for the following morning; the first on Oilseed Processing, and the second on Thin Layer Chromatography of Fatty Materials. Two plant trips are also scheduled.

Members of the Southwest Section have been invited to attend the Fresno meeting.

Southwest Section

The Southwest Section of the American Oil Chemists' Society held its first meeting of the current year at the Rodger Young Auditorium in Los Angeles, September 13. The speaker for the evening was R. T. Doughtie, of the A.O.C.S. Examining Board, who discussed the role that Referee Chemists play in the American Oil Chemists' Society activities.

• *New Members*

Active

Jerzy J. Biezenski, Research Associate, Maimonides Hospital, Brooklyn, N. Y.

Fred E. Boettner, Senior Research Chemist, Rohm and Haas Co., Bristol, Pa.

Kenneth M. Brobst, Head, Analytical Laboratory, A. E. Staley Mfg. Co., Decatur, Ill.

Colin J. Cox, Development Engineer, Lever Brothers Ltd., Toronto, Ont., Canada

John T. Grogan, Product Development Chemist, Standard Brands, Inc., Indianapolis, Ind.

Walter G. Jennings, Associate Professor and Associate Chemist, University of California, Davis, Calif.

Kimitoshi Nakazawa, Senior Managing Director, Japan Margarine & Shortening Makers Association, Tokyo, Japan.

Robert J. McPherson, Assistant Chemist, Kershaw Oil Mill, Kershaw, S. C.

Robert L. Moore, Jr., American Mineral Spirits Div. of The Pure Oil Co., Atlanta, Ga.

Myron A. Schmutzer, Manager, Chemical Division, United States Testing Co., Inc., Hoboken, N. J.

Marlyn W. Stephens, Chemist, Analytical Research Section, Emery Industries, Inc., Cincinnati, O.

Joseph R. Swartwout, Associate Professor, Dept. of Ob-Gyn., Emory University School of Medicine, Atlanta, Ga.

Bob W. Wright, Assistant Manager, Chemical Division, United States Testing Co., Inc., Hoboken, N. J.

Active Junior

P. R. Lakshmanan, Graduate Student, Dept. of Coatings, North Dakota State University, Fargo, N. D.

forming solutions and true liquids at concentrations above 2 to 10% in water, and (2) an alkali metal sulfonate of mixed hydrocarbons substantially comprising alpha and beta methyl naphthalenes and dimethyl naphthalenes, the alkali metal sulfonate being essentially free from olefins and constituting from 40 to 70% of the mixture; (b) from 30 to 38 parts of a water-soluble polyphosphate, and (c) from 20 to 30 parts of an ammonium sulfate. The composition will form water solutions at all concentrations up to about 48% of solids at normal temperatures.

SOAP BAR COMPOSITIONS. W. A. Kelly (Lever Brothers Co.). *U. S. 3,043,778.* A bar of soap which disperses soap from scum in hard water consists of (a) 5–25% by weight of an alkali metal salt of a fatty acid methyl taurate, the fatty acid containing from 12 to 20 carbon atoms and the cationic portion being sodium or potassium; (b) 15–60% of a water soluble alkali metal coconut soap, and (c) 2–25% of free fatty acid containing 12 to 20 carbon atoms. The bar of soap may also contain from 0.1 to 1.0% of 3,4',5-tribromosalicylanilide. It has a pH, as measured in a 10% aqueous solution, of 8.6 to 9.4.

DETERGENT TABLETS. J. P. Parke and D. D. Penketh (Lever Brothers Co.). *U. S. 3,043,779.* Described is a detergent composition in the form of a hard coherent tablet which does not become soft and slimy or sticky on being used or stored after use and which does not dissolve with undesirable rapidity in warm water. It contains from 55–85% of a water-soluble alkali metal soap, from 15–30% of a water-soluble alkali metal fatty-acyl-aminomethane sulphonate in which the fatty-acyl radical has from 8–18 carbon atoms, and from 0.5–10% of a water-soluble salt of a monocarboxylic or monosulphonic acid having from 1 to 8 carbon atoms in the molecule.

DETERGENT COMPOSITION CONTAINING A 3,5-DIAMINOPYRAZINE-2,6-DICARBOXYLIC ACID DERIVATIVE. A. F. Daglish, R. Vonderwahl, and G. A. Tillotston (J. R. Geigy A.-G.). *U. S. 3,043,780.* The described composition consists of a water soluble, synthetic, organic anion-active detergent (a water soluble soap or synthetic, organic non-soap anionic sulphated detergent) and from 0.001 to 5% by weight of a 3,5-bis-diamino-pyrazine-2,6-dicarboxylic acid diamide.

Index to Advertisers

PRODUCTION OF ETHYLENE GLYCOL, PROPYLENE GLYCOL & PENTAERYTHRITOL
(Millions of Pounds)

	Ethylene Glycol	Propylene Glycol	Penta-erythritol	Total**
Year:				
1957	1,199.9	98.4	56.5	1,354.9
1958	1,145.5	42.0 [b]	53.3	1,240.8
1959	1,214.6 [a]	151.5 [a]	64.1 [a]	1,430.2
1960	1,297.3 [c]	152.0 [c]	64.3 [c]	1,513.5
1961	1,142.1 [d]	143.3 [d]	62.4 [d]	1,347.7
1960				
May	97.1	....	5.1	102.2
June	100.6	15.1	5.3	121.0
July	112.6	13.1	4.0	129.7
August	121.5	12.1	5.7	139.3
September	115.6	9.1	5.6	130.3
October	111.7	13.7	5.6	131.0
November	104.9	8.7	5.2	118.8
December	108.8	....	5.1	113.8
1961:				
January	119.5	12.0	5.5	137.0
February	101.0	12.7	4.4	118.1
March	101.3	14.0	5.3	120.7
April	95.7	14.1	5.4	115.3
May	98.4	15.4	5.3	119.1
June	97.0	15.1	5.1	117.2
July	94.4	16.0	4.6	115.1
August	99.3	15.3	5.0	119.7
September	87.7	10.7	5.2	103.6
October	97.5	12.3	5.9	115.6
November	95.2	10.8	4.9	110.9
December	97.3	11.9	5.1	114.3
1962:				
January	91.3	19.1	5.8	116.1
February	80.8	15.4	4.5	100.7
March	87.9	16.0	4.7	108.5
April	88.8	12.2	5.0	106.0
May	98.4	17.2	4.8	120.4
June	103.7	13.6	4.5	121.7

** Totals may not agree exactly because of independent rounding of figures.

[a] Revised, but does not agree with total monthly figures.

[b] Incomplete—sum of five monthly figures only. The Tariff Commission did not publish an official 1958 annual figure for propylene glycol.

[c] Official figure, but does not agree with total of the months.

[d] Revised, but still preliminary. Totals shown do not agree with total of monthly figures.

SOURCE: U. S. Tariff Commission